Chemistry 102 – GENERAL CHEMISTRY II

Day/Time: MW 5:15-6:40 pm **Section: 3157 & 3158** **Room: CMS 236**

I'm Professor
Said Pazirandeh
Welcome to my class!

Office Phone/Voicemail:
(818) 364-7705

Office Location:
Center for Math & Sci (CMS) Room #242
Office Hours:
TuW 3:30-4:30, or another more convenient time for both of us.
Drop by for help or just to say hi!

You can always email
profpaz@earthlink.net
paziras@lamission.edu

Highly Recommended:

- Attend every class
- Bring lecture notes
- Annotate notes
- Participate in discussions
- Put deadlines in your calendar
- Plan ahead
- Complete all assignments

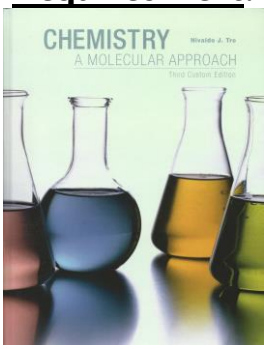
SUCCESSFUL STUDENTS ARE PREPARED, ON TIME, FOCUSED, AND RESPECTFUL

WEBSITE FOR SUCCESS

WWW.PROFPAZ.COM
This site has digital resources:
⇒class syllabus
⇒ practice problems
⇒study guides
⇒ practice exams
⇒ lab videos

HANDOUTS on class website for LECTURE OUTLINE. PRINT- bring to every class!

Required Text:



Chemistry: A Molecular Approach
Nivaldo Tro, 3rd Ed.
Custom published for LAMC
(ISBN 9781269342506)

LOS ANGELES MISSION COLLEGE-SPRING 2014
CHEMISTRY 102-SECTIONS 3157 & 3158
Lecture: MW 5:15-6:40 ; Room: CMS-236
Lab (3157): MW 6:50-10:00 ; Room: CMS-201
Lab (3158): MW 6:50-10:00 ; Room: CMS-203

INSTRUCTOR: Said Pazirandeh — **OFFICE PHONE: (818)364-7705**
E-MAIL: profpaz@earthlink.net — **OFFICE: CMS 242**
WEBSITE: www.profpaz.com — **OFFICE HOURS: TW 3:30-4:30**

INSTRUCTOR (Lab): Maria Fenyes — **OFFICE PHONE: (818) 364-7600 X-4336**
E-MAIL: mariafenyes@earthlink.net — **OFFICE: CMS 201**
WEBSITE: www.proffenyes.com — **OFFICE HOURS: MW 4:30 – 5:15 p.m.**

INSTRUCTOR (Lab): Charles Mallory — **OFFICE PHONE: (818) 364-4223**
E-MAIL: mallorcp@lamission.edu — **OFFICE: TBA**
WEBSITE: http://www.themalloryfamily.net/ — **OFFICE HOURS: TBA**

1. **PREREQUISITES**:
   - Chemistry 101 with a grade of C or better.
2. **RECOMMENDED**:
   - Completion of Math 260 (Precalculus) or a higher level Math class with a grade of "C" or better

3. **TEXTBOOK**:
   - Required: "**Chemistry: A Molecular Approach**", Nivaldo Tro (3rd Edition; ISBN 978-1-269-34250-6)
   - Copy of the Textbook will be available on Reserve in the Library.
   - See helpful hints for selecting textbook available on profpaz.com.

4. **LABORATORY NOTEBOOK:**
   - Required: This is a **quadrille paper,** hard cover "Comp Book", available in the LAMC Bookstore and in general office supply stores. You are required to have your laboratory notebook by the 2nd class meeting.
   - You are required to report all laboratory work in your Laboratory Notebook (See Page 5 f this outline for the proper use of the Laboratory Notebook).

5. **SCIENTIFIC CALCULATOR**
   - Need not to be an expensive type, but it must perform the following operations: Addition, Subtraction, Multiplication, Division, Square Root, 1/x, and Logarithms.
   - You are required to have your calculator with you for all class sessions (lectures and labs).

6. **SAFETY GOGGLES**
   - Unless specifically instructed otherwise by your instructor, you must wear safety goggles during laboratory work. Safety goggles are available for purchase in the LAMC Bookstore. You are required to have your safety goggles by the second class session. You may keep your goggles locked in your laboratory locker.
   - **While in the laboratory, students must wear safety goggles at all times, unless otherwise directed by the instructor. Failure to wear goggles unless directed by the instructor is grounds for dismissal from the laboratory.**

7. **PERIODIC TABLE OF THE ELEMENTS**
   - You are required to have your own Periodic Table of the Elements with you, for all class sessions.
   - The particular type of Periodic Table used for this course is available online at my website.

STUDENT LEARNING OUTCOMES

1. Describe, explain and model chemical and physical processes qualitatively at the molecular level in order to explain macroscopic properties.
2. Solve quantitative chemistry problems through integration of multiple ideas and demonstrate reasoning clearly and completely.
3. Analyze results of laboratory experiments, evaluate sources of error and prepare clear and organized laboratory reports.
4. Perform laboratory techniques safely and accurately and maintain a laboratory notebook according to standard scientific guidelines.
5. Design, construct and interpret graphs accurately.

GRADING SCALE

- You will be assigned a unique student code and password and can check your grade online. More detailed information will be given by the instructor after the 2nd week of class.

- Your grade in the class is composed of the following components:

ASSIGNMENT	*POINTS*	*% OF TOTAL*
• **QUIZZES**	150	**15**
• **TESTS** (3 x 100 points each)	300	**30**
• **FINAL EXAM**	150	**15**
• **LABORATORY REPORTS & UNKNOWNS**	200	**20**
• **LABORATORY EXAMS**	190	**19**
• **LABORATORY PARTICIPATION**	10	**1**
TOTAL	1000	**100**

- The grading scale in the class is as follows:

A	90% – 100%
B	80% – 90%
C	65% – 80%
D	55% – 65%
F	less than 55%

TENTATIVE LECTURE SCHEDULE

Week	Date	Text Reference	Topic
1	Feb. 10	13.1-13.2	Introduction to class – Rates of Reactions
	Feb. 12	13.3-13.4	Rate Laws
2	Feb. 17	------	President's Day (college closed)
	Feb. 19	13.5	Temperature & Rate
3	Feb. 24	13.6-13.7	Reaction Mechanisms & Catalysis
	Feb. 26	14.1-14.5	Chemical Equilibrium & Equilibrium Constant
4	Mar. 3	14.6-14.8	Calculations with Equilibrium Constant
	Mar. 5	14.9	Le Chaterlier's Principle/Review for Test 1
5	**Mar. 10**	-------	**Test 1 (Chapters 13–14)**
	Mar. 12	15.1-15.4	Introduction to Acids & Bases
6	Mar. 17	15.5-15.6	Calculating pH of Strong Acids
	Mar. 19	15.6-15.7	Weak Acid & Base Equilibria
7	Mar. 24	15.9-15.11	Acid Strength & Molecular Structure
	Mar. 26	15.8	Acid-Base Properties of Salts
8	Mar. 31	-------	Cesar Chavez Day (college closed)
	April 2	-------	**Test 2 (Chapters 15)**
9	April 7-13	------	Spring Break (College closed)
10	April 14	16.1-16.3	Buffers/Common-Ion Effect
	April 16	16.4	Acid-Base Titrations
11	April 21	16.5	Solubility Equilibria
	April 23	16.8	Complex-Ion Equilibria
12	April 28	16.6-16.7	Selective Precipitation/Qualitative Analysis
	April 30	--------	**Test 3 (Chapters 16)**
13	May 5	17.1-17.4	Laws of Thermodynamics
	May 7 **May 3**	17.5-17.8 -------	Gibbs Free Energy & Spontaneity ***Last day to drop with a "W" (In Person)***
14	May 12	17.9.	Free Energy & Equilibrium
	May 14	18.2	Balancing Redox Equations
15	May 19	18.3	Voltaic Cells & Cell Notations
	May 21	18.4	Standard Cell Potential/Free Energy & Equilibrium
16	May 26	------	Memorial Day (college closed)
	May 28	18.6	Cell Potential and Concentration
17	June 2	------	Review for Final
	June 4 (5:30-7:30)		**FINAL EXAM (Chapters 17, 18)**